ANSWER KEY

Indian National Junior Science Olympiad

INJSO 2025

Que.	1	2	3	4	5	6	7	8	9	10
Ans.	D	B	C	A	A	D	C	C	C	B
Que.	11	12	13	14	15	16	17	18	19	20
Ans.	C	B	B	A	A	B,C,D	A,B	B,D	B,D	A,D
Que.	21	22	23	24						
Ans.	B,D	B,C	B,D	A,B,D						

Indian National Junior Science Olympiad-2025

Date of Examination : 01st February, 2025

SOLUTIONS

SECTION-I

1. An unusual and fatal disease was observed in a sheep herd. Treatment with a broad-specrum antibiotic proved ineffective. Upon autopsy, veternarians discovered extensive damage and deformity in the brain tissue. Noting similarities to mad cow disease. they sought to confirm their hypothesis by processing the damaged tissue. The tissue was then divided, blended in a mixer, and filtered through a 0.2 urn sieve. The resulting filtrate was then divided into three fractions for further testing under different treatments:

Fraction X: Exposed to 90°C for 20 minutes.

Fraction Y: Treated with high concentration of sodium hypochlorite for 60 minutes.

Fraction Z: Treated with highly active nuclease for 20 minutes at 25°C.

If these fractions are injected into healthy individual sheep, which of these fractions will cause same disease in sheep?

(A) Only fraction X
(B) Only fraction Y
(C) Fractions Y and Z
(D) Fractions X and Z

Ans. **(D)**

Sol. Fractions X and Z

2. A farmer in Udupi, a town in North costal Karnataka, decided to boost her income by planting popular fruit trees like Alphonso mangoes and Kashmiri apples in her very own optimised greenhouse facility. She purchased fresh fruits from organic farms of Goa and Kashmir, respectively. The seeds were removed and sowed in appropriate soil conditions and waited for them to germinate. Seeds of mango germinated easily whereas apple seeds remained intact and unsprouted. What is the likely reason for lack of germination from the apple seeds?

(A) The warm temperature but not humidity was required to break dormancy of apple seeds.
(B) Seeds needed cold period to break its dormancy.
(C) Seeds came from a farm in a colder region rather than a warmer region.
(D) Apple seeds have amygdalin, which upon metabolism releases cyanide, deters breaking of dormancy.

Ans. **(B)**

Sol. Seeds needs cold period to break its

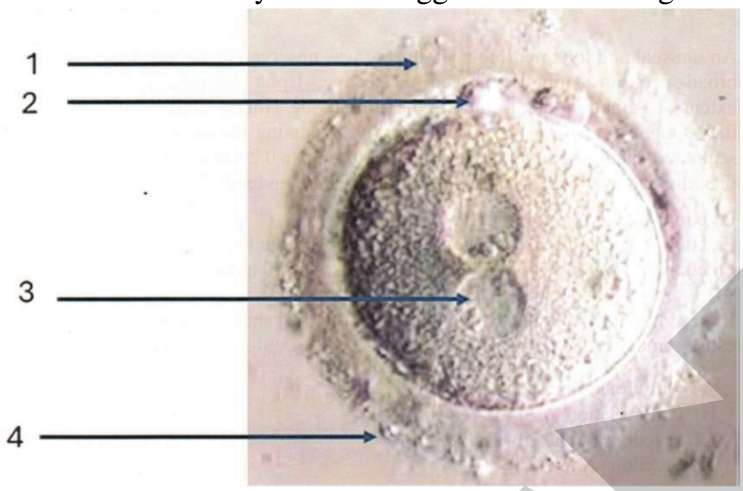
3. Two hundred years ago, Assam (India) had a large population of migratory stork birds. During that time, infections caused by the West Nile Virus (WNV1 serotype) resulted in a mortality rate of only 10%. However, in the early 1900s, widespread superstitious beliefs and habitat destruction led to the mass killing of storks and the destruction of their nests, causing a drastic decline in their population. Thanks to decades of conservation efforts and increased awareness, the stork population has rebounded to a significant size. However, a recent outbreak of WNV1 has shown a 40% mortality rate among the birds. Which of the following option rightly explains the higher mortality rate?

(A) Natural selection
(B) Chance event
(C) Genetic drift
(D) Migratory effect

Ans. **(C)**

Sol. Genetic drift

4. Identify the components of the freshly fertilized egg and select the right labeling options.



(A) 1) Zona pellucida; 2) Polar body; 3) Pronucleus; 4) Sperm
(B) 1) Pronucleus; 2) Polar body; 3) Zona pellucida; 4) Sperm
(C) 1) Zona pellucida; 2) Sperm; 3) Polar body; 4) Zona pellucida
(D) 1) Sperm; 2) Pronucleus; 3) Polar body; 4) Zona pellucida

Ans. **(A)**

Sol. 1) Zona pellucida; 2) Polar body; 3) Pronucleus; 4) Sperm

5. The following property of nucleic acid is most relevant for its ability to function as genetic material when compared to a protein.
(A) Complementary base pairing.
(B) Having a high density of negative charge enables them to form chromatins.
(C) Having 4 (bases) instead of 20 (amino acids) in numbers.
(D) Creates maximum structural diversity.

Ans. **(A)**

Sol. Complementary base pairing is unique properties of nucleic acid.

6. In his examination, Aryan writes differences between voltaic cell and electrolytic cell. But he did not get full marks in this answer as he wrongly placed his statements in wrong column. Find ALL INCORRECT difference(s) from his answer given below:

Differences	Voltaic cell	Electrolytic cell
(i)	Anode is the positive electrode, and cathode is the negative electrode.	Anode is the negative electrode, and cathode is the positive electrode.
(ii)	The current flows in the cell from negative to positive electrode.	The current flows in the cell from positive to negative electrode.
(iii)	A spontaneous chemical reaction occurs in the cell to produce electricity.	A non spontaneous chemical reaction takes place in the cell.
(iv)	The external battery drives the electrons that flow through	The electrons flow from anode to cathode through the

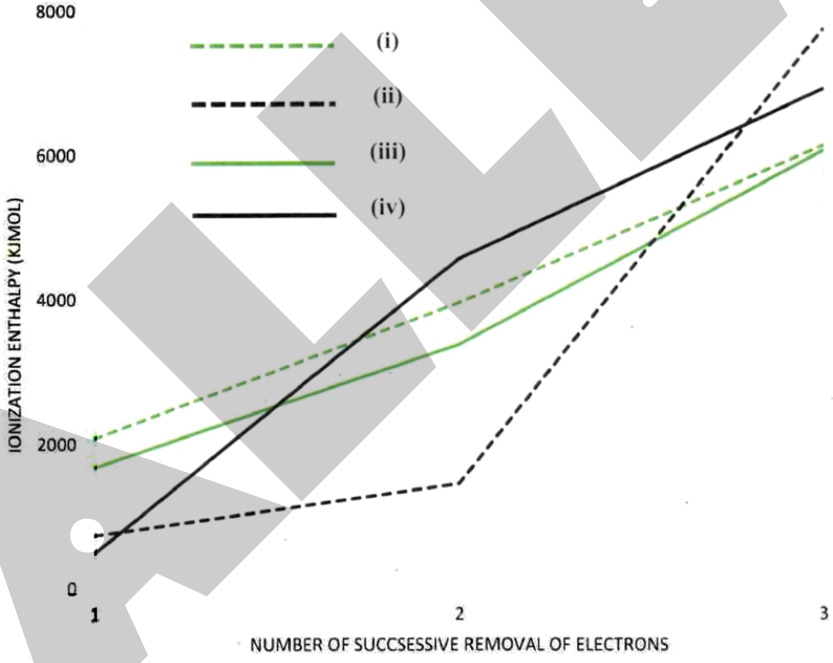
	cathode to anode.	external circuit.
(v)	Oxidation takes place at positive electrode.	Oxidation takes place at negative electrode.
(vi)	Reduction takes place at positive electrode.	Reduction takes place at negative electrode.
(vii)	Discharge of ions occurs at both electrodes.	Discharge of ions occur only at cathode.

(A) (vii)
(B) (i), (iv), (v)
(C) (iv), (vii)
(D) (i), (iv), (v), (vii)

Ans. (D)

Sol. In voltaic cell cathode is positive electrode at which reduction occurs and anode is negative electrode at which oxidation occurs. Current flows from positive electrode to negative electrode in external circuit.

7. In the following graph, ionization enthalpies of successive elements in the periodic table are plotted against number of successive removal of electrons. Analyse the graph and predict the elements from the graph.



(A) (i)-N, (ii)-O, (iii)-C, (iv) - B
(B) (i)-N, (ii)-C, (iii)-Be. (iv)-B
(C) (i)- Ne, (ii)- Mg, (iii) - F, (iv)-Na
(D) (i)-P, (ii)-S, (iii)-Si, (iv)-AI

Ans. (C)

Sol. (i) – Ne
(ii) – Mg
(iii) – F
(iv) – Na

8. In the Haber process. 60 litres of hydrogen gas and 20 litres of nitrogen gas were taken for reaction which yielded only 25% of the expected product due to malfunctioning of the production unit. What is the mole fraction of ammonia in the mixture and what is the volume of ammonia extracted from the mixture at the same pressure?

(A) 0.33, 40L (B) 0.14, 20L (C) 0.14, 10L (D) 0.33, 10L

Ans. (C)

Sol. $\underset{15\text{lt}}{\overset{20\text{ lt}}{N_{2(g)}}} + \underset{15\text{lt}}{\overset{20\text{ lt}}{3H_{2(g)}}} \longrightarrow \underset{40\times\frac{25}{100}=10\text{ lt}}{\overset{20\text{ lt}}{2NH_{3(g)}}}$

Mole fraction of $NH_3 = \frac{10}{70} = 0.14$

9. Given below are the boiling points of hydrocarbons containing 6 carbons. Choose the correct statement based on the data given.

neohexane	2,3-dimethylbutane	isohexane	n-hexane	cyclohexane
50°C	58°C	60°C	69°C	81°C

(A)Cyclohexane has a higher boiling point than n-hexane because cyclohexane has a higher molecular weight than n-hexane.

(B) n-hexane has a higher boiling point than isohexane, because n-hexane is a linear molecule and has lesser mtermolecular Van der Waal's interaction than isohexane.

(C) Replacing a carbon in any of the above given 6-carbon compounds with an oxygen result in a molecule whose boiling point will be higher than the parent compound.

(D) Fractional distillation is an efficient way to separate every hydrocarbon from each other from a mixture that contains equal amounts of all these compounds.

Ans. (C)

Sol. Due to replacement of carbon by oxygen molecular weight and polarity of molecule increase, So intermolecular attraction among them increase and So, their boiling point increase.

10. During the Second World War, engineers developed rocket-powered aircraft. The air craft carried these two liquids: hydrazine (N_2H_4) and hydrogen peroxide, (H_2O_2). When these two liquid mix in the combustion chamber, they evaporate and then react rapidly to form nitrogen gas and steam. The equation for the reaction is

$$N_2H_4 + 2H_2O_2 \rightarrow N_2 + 4H_2O$$

Bond dissociation energy (BDE) is the amount of energy required to break a specific chemical bond in a molecule in its gaseous state.

BOND	BOND DISSOCIATION ENERGY in kJ/mol
N≡N	945
N−N	159
N−H	391
O−O	143
O−H	463

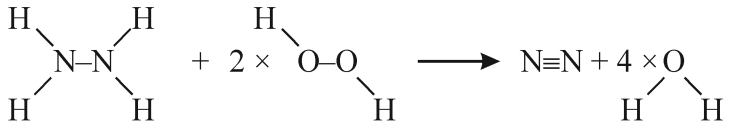
Use the data given in the table above to calculate the total amount of energy released or absorbed by the chemical process.

(A) Absorbs 788 kJ/mol (B) Releases 788 kJ/mol

(C) Absorbs 2640 kJ/mol (D) Releases 2640 kJ/mol

Ans. (B)

Sol.



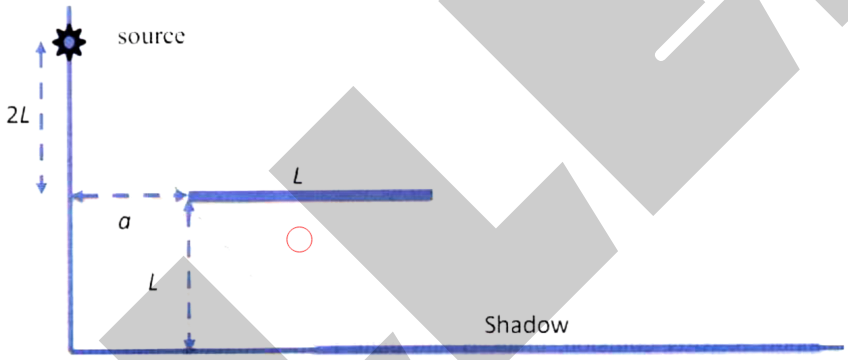
$\Delta H = B.E._{Reactant} - B.E._{Product}$

$= [(4 \times 391) + (1 \times 159) + 2 \times \{(2 \times 463) + (1 \times 143)\}] - [(1 \times 945) + (8 \times 463)]$

$= 3861 - 4649$

$= -788 kJ/mole$

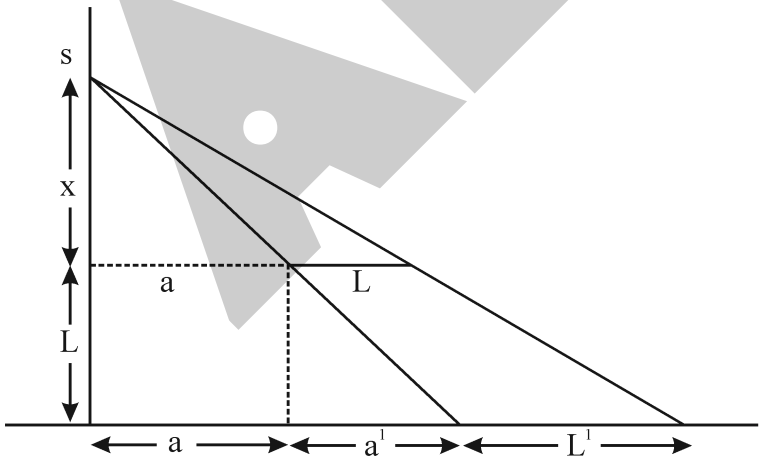
11. The shadow of a plank of length L is formed on the floor due to a point shown in the figure. If the source comes down by 20% of the present vertical length (2L) from the plank, find the percentage change in the length of the shadow formed at the previous level.



(A) 25% (B) 6.33% (C) 8.33% (D) 21%

Ans. (C)

Sol.



By smiliarity of triangles

$$\frac{x}{a} = \frac{L}{a_1} \Rightarrow a^1 = \frac{aL}{x} \qquad (1)$$

For previous position $\rightarrow a^1 = \frac{a \times L}{2L} = \frac{a}{2}$ (2)

For new position $\rightarrow a^1 = \frac{aL}{\frac{8L}{5}} = \frac{5a}{8}$ (3)

By similiarity of triangles $\rightarrow \frac{a+L}{x} = \frac{a+a^1+L^1}{x+L}$

For previous position $\rightarrow \frac{a+L}{2L} = \frac{a+\frac{a}{2}+L^1}{2L+L}$

$L^1 = \frac{3L}{2}$ (4)

For new position $\rightarrow \frac{a+L}{\left(\frac{8L}{5}\right)} = \frac{a+\frac{5a}{8}+L^1}{\frac{8L}{5}+L} \Rightarrow L^1 = \frac{13L}{8}$

Percentage change $= \frac{\frac{13}{8}-\frac{3}{2}}{\frac{3}{2}} \times 100 = 8.3\%$

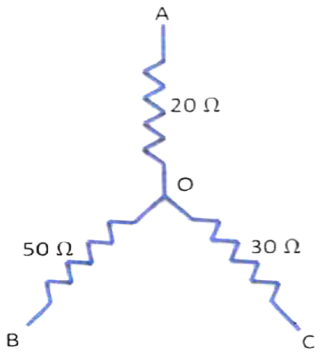
12. A portion of an electrical network is shown in the adjacent figure. Electrostatic potential of terminals A and B are V_A = 150 V and V_B = 40 V respectively, the terminal C is open. Read the following statements:

Statement S_1: If the terminal C is at potential V_C = 30 V, the current passing through 20 Ω resistor will be 3 A.

Statement S_2: If terminal C is connected to the ground (i.e. V_C = 0 V), no current will How through the 50 H resistor.

Statement S_3: If terminal C is connected to terminal A, the electric power dissipated in the 30 Ω resistor will be about 15.1 W.

Now, choose the correct option.



(A) Statements S_1. S_2 and S_3 are incorrect.

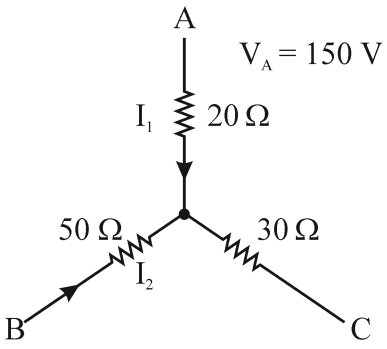
(B) Statements S_1 and S_3 are correct, and S_2 is incorrect.

(C) Statements S_2 and S_3 are correct, and S_1 is incorrect.

(D) All the statements S_1, S_2 and S_3 are correct.

Ans. (B)

Sol.



$V_B = 40$ V

If $V_C = 30$ V $\rightarrow$ By KCL at (0) $\rightarrow$

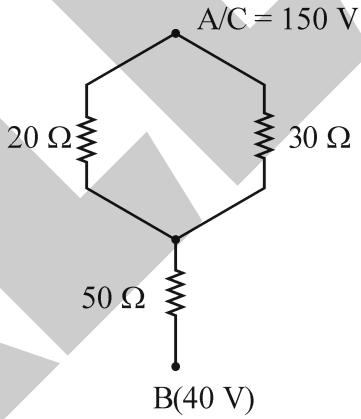
$V_0 = 90$ volt

$I_1 = \frac{150-90}{20} = \frac{60}{20} = 3A$

If $V_C = 0$ V then $V_0 = 80.32$ volt

$I_2 \neq 0$

If $V_A = V_C \rightarrow$



$V_A - V_O = 21.3$ volt

$P_{30} = \frac{(21.3)^2}{30} = 15.123$ W

13. A beaker partly filled with a liquid of density 1.2 g cm^{-3} is placed on a digital balance, and the reading is 440 g. A metal piece is tied at the end of a string, and the free end is held by hand. The metal piece is now completely immersed in the liquid without touching any side or bottom of the beaker. The reading in the digital balance has now become 500 g. The string is cutoff, and the metal-piece is dropped into the liquid without spilling water. It gets completely immersed in the liquid, rests at the bottom, and the reading becomes 680 g. The density of the metal is

(A) 5200 kg m^{-3} (B) 4800 kg m^{-3} (C) 4000 kg m^{-3} (D) 6000 kg m^{-3}

Ans. (B)

Sol. 0.500 g = 0.440 g + B

B = Buoyant force on metal body

g = acceleration due to gravity

$\Rightarrow$ B = 0.06 g …(1)

$\Rightarrow$ $B = P_L v_m g$ …(2)

$\Rightarrow$ P_L = density of liquid = 1.2 g/cm^3

$\Rightarrow$ V_m = volume of metal body

$\Rightarrow$ $0.680 \times g = 0.440 \times g + P_m \ V_m.g$

$\Rightarrow$ $P_m V_m = 0.24$ ....(3)

$\Rightarrow$ From (1) and (2)

$\Rightarrow$ $V_m = 5 \times 10^{-5}$ m^3

$\Rightarrow$ From (3)

$\Rightarrow$ $P_m = 4800$ kg/m^3

$\Rightarrow$ Ans = B

14. A small block B of mass 0.6 kg starts moving on a rough horizontal surface with an initial velocity of 6.0 m s^{-1}. The coefficient of friction between the block and surface. $\mu = 0.1$. frictional force on the block is expressed as

$f = \mu \times$ Normal force between the block and surface.

After travelling through a distance of 10.2 m on the rough horizontal surface, block B collides with a plate P of mass 0.4 kg connected to the free end of an elastic massless spring of the force constant k = 400 N m"1. The other end of the spring is attached to a rigid support S. After collision, block B and plate P stick together and move on a perfectly smooth surface ahead to compress the spring. Assume that the spring obeys Hooke's law during compression. There is no frictional force on the block during the collision and thereafter. You may neglect the size of the block B and plate P for calculation.

Statement 1: The velocity of block B just before the collision is close to 4.0 m s^{-1}.

Statement 2: Maximum compression produced in the spring is 12 cm.

Statement 3: Work done by the friction on the block during the entire motion is close to –6.0 J.

Choose the correct option.

(A) Statements 1, 2 and 3 arc correct.

(B) Statements 1 and 3 alone are correct.

(C) Statements 1 and 2 alone are correct.

(D) Statements 2 and 3 alone are correct.

Ans. **(A)**

Sol. Let's say the mass of B = m_1 = 0.6

Let's say the mass of P = m_2 = 0.4

$\Rightarrow$ Work done by friction = W_{fr}

$= -\mu m_1 g \times d$

$= -0.1 \times 0.6 \times 10 \times 10.2$

$= -6.12\ J \simeq -6\ J$

$\Rightarrow$ Statement 3 is correct.

$\Rightarrow$ Applying work energy theorem on block before collision:

$\Rightarrow -\mu m_1 g\, d = \frac{1}{2} m_1 v_1^2 - \frac{1}{2} m_1 \times 6^2$

$\Rightarrow -6.12 = \frac{1}{2} \times 0.6 \times v_1^2 - \frac{1}{2} \times 0.6 \times 6^2$

$\Rightarrow v_1 \simeq 4$ m/s

$\Rightarrow$ Statement 1 is correct.

$\Rightarrow$ Applying momentum conservation during collision:

$\Rightarrow m_1 \times v_1 = (m_1 + m_2)V_f$, $V_f = 2.4$ m/s

Applying energy conservation to find maximum compression in spring:

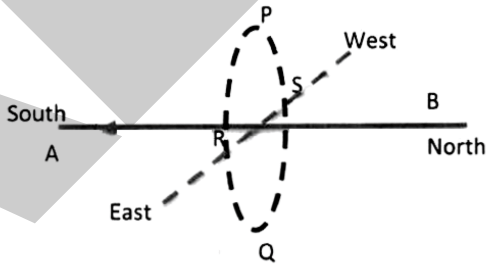
$\Rightarrow \frac{1}{2} K [0^2 - X^2_{max}] = 0 - \frac{1}{2} (m_1 + m_2)V_f^2$

$\Rightarrow \frac{1}{2} \times 400\ [0^2 - X^2_{max}] = -\frac{1}{2} \times 1 \times (2.4)^2$

$\Rightarrow X_{max} = 12$ cm

$\Rightarrow$ Statement 2 is correct.

15. AB is a section of a long conductor carrying constant current (indicated by the arrow) Placed along the north-south direction, as shown in the figure P, Q, R and S are four points around the conductor near the centre and close to the conductor as shown in the figure. A small magnetic compass can be placed horizontally at point P and Q. At point R and S, the plane of compass is placed vertically (axis along Fast-West). The magnetic needle is free to rotate in the plane of the compass, and its mass is negligible. The direction of the north pole of the compass deflects in the direction of the resultant magnetic field at that point.



Choose the correct statement.

(A) At point Q, the North pole of the compass will deflect towards the But,

(B) At point P, the North pole of the compass will deflect towards the East.

(C) At point R, the North pole of the compass will deflect vertically down.

(D) At point S, the North pole of the compass will not deflect.

Ans. **(A)**

Sol. (A) At point Q, North pole will point towards East.

(B) At point P, North will point towards West.

(C) At R, vertically up

(D) At S, vertically down

SECTION-II

16. Scientists have discovered that the desert moss *Syntrichia caninervis*, found in Antarctica, can survive extreme Martian conditions. Known for its resilience to drought, further research shows it can withstand temperatures as low as -196°C, low oxygen (95% CO_2), and tolerate high levels of gamma radiation, making it a strong candidate for colonizing Mars. Researchers suggest it could play a crucial role in building biologically sustainable habitats for humans on other planets. Which of the following properties makes S. *Caninervis* a potential candidate for colonizing Mars?

(A) Cellular activities generate water as a net by-product.

(B) High reproductive rate as an advantageous trait.

(C) Sexual reproduction.

(D) Low nutrient requirements.

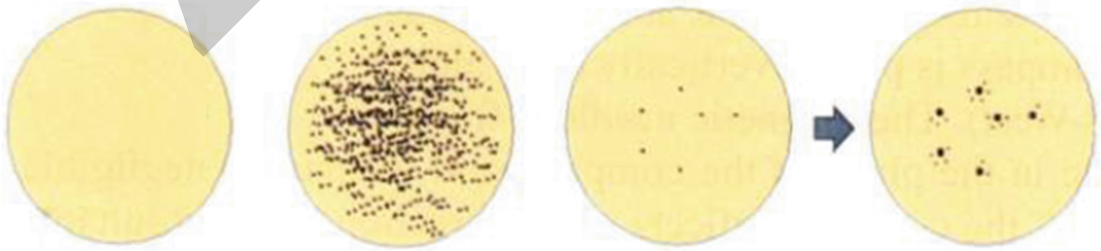
Ans. **(B, D, C)**

Sol. *

17. Bacterial transformation is process where a bacterial cell takes up DNA from its environment. Gram-positive bacteria have this as an inherent property, but Gram-negative bacteria do require assistance. A student is tasked to perform bacterial transformation of an Escherichia coli strain with a suitable expression plasmid that allows the student to select the transformed cells having ampicillin resistance. The resistance to ampicillin was due to its degradation by the encoded protein that is secreted outside the cell. After transformation procedure, the bacterial cells were spread on plate containing ampicillin and incubated for 12 and 24 hr at 37°C. Parallelly, two more spread plates were done. One, to be sure of quality of antibiotic, bacterial cells alone were spread on plate with antibiotics.

Second, to be assured that the bacterial strain is viable, they were spread on another plate without any antibiotics. These control plates were incubated as mentioned above. The experimental observations are labelled as Plate 1, 2, 3a and 3b where the black dots on the plate indicate growth of bacterial colonies. The proposed hypothesis that can potentially explain the observations are stated as the options below. Mark the statements which is are true.

Components	Plate 1	Plate 2	Plate 3	
Bacterial cells spread	10^6	10^6	10^6	
Plate with ampicillin	100 mg/ml	0 mg/ml	100 mg/ml	
Plasmid DNA	0 ng	10 ng	10 ng	
Incubation time	12 hr	12 hr	**Plate 3a**: 12 hr	**Plate 3b**: 24 hr



(A) Instead of ampicillin, if a monobactam was used (both belong to same family of classification), a similar profile is expected to be observed.

(B) The formation of big and small colonies as observed in Plate 3b results from fast growing and slow-growing transformants, usually observed after longer periods of incubation.

(C) If Plate 1 was incubated further for 24 hrs, some small colonies would have appeared like in Plate 3b.

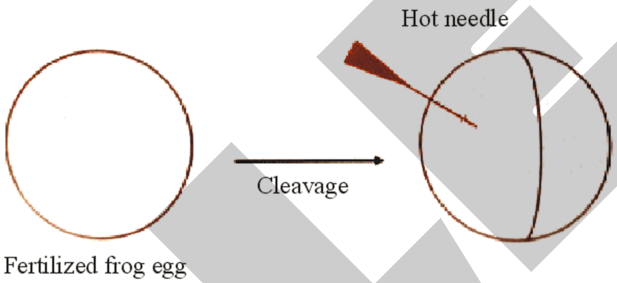
(D) In Plate 3b. longer incubation led to ampicillin depletion around the transformant colonises and that allowed a few non-transformed cells to grow around as tiny colonies.

Ans. (A, B)

Sol. *

18. Imagine a fascinating scenario where a fertilized frog egg begins its journey of development. This zygote has now undergone its very first division, resulting in two identical cells that together hold the potential to create a complete organism. This is a crucial phase in the early development of the frog embryo, where each of these two cells is identical in appearance. A scientist thought what would happen if one of the cells is pricked with a hot needle, without damaging the other cell?

Mark the statements which is/are true.



(A) The Other cell develops into half embryo Inn then doesn't survive.1

(B) The other cell develops into complete embryo.

(C) The other cell develops into a muture frog of half the natural size.

(D) The pricked cell does not contribute to the development of embryo.

Ans. (B, D)

Sol. (B) The other cell develops into complete embryo.

(D) The pricked cell does not contribute to the development of embryo.

19. In the final examination of Class X, a chemistry teacher asked her students to make different solutions of acid and alkali. The students prepared different mixtures according to the table given below. Each student ultimately mixed the solutions that they prepared into one flask.

Student	$HCl_{(aq)}$ solution (mL)			$H_2SO_{4(aq)}$ solution (mL)			$NaOH_{(aq)}$ solution (mL)			Pure water (mL)
	0.1 M	0.2 M	0.5 M	0.1 M	0.2 M	0.5 M	0.1 M	0.2 M	0.5 M	
U	-	20	-	25	-	-	-	-	20	35
V	50	-	-	-	-	10	-	25	-	15
W	25	-	-	-	25	-	-	-	25	25
X	-	20	-	-	20	-	20	-	-	40
Y	-	-	14	35	-	-	-	-	8	43
Z	-	-	20	-	15	-	-	30	-	35

After the mixing activity, which set of the students got a solution of approximately pH = 1?

(A) U, W, X (B) V, X, Z (C) U, W, Y (D) V, Y, Z

Ans. (B, D)

Sol. Student U : $[H^+]$ or $[OH^-] = \frac{(0.2\times20)+(0.2\times25)-(0.5\times20)}{(20+25+20+35)} = \frac{-1}{100} = -10^2$

$\therefore$ $[OH^-] = 10^{-2}$ M

$\therefore$ pOH = 2 & pH = 12

Student V : $[H^+]$ or $[OH^-] = \frac{(0.1\times50)+(1\times10)-(0.2\times25)}{50+10+25+15}$

$\therefore$ $[H^+] = \frac{12.5}{100}$M

$\therefore$ pH = 0.9

Student W : $[H^+]$ or $[OH^-] = \frac{(0.1\times25)+(0.4\times25)-(0.5\times25)}{25+25+25+25} = 0$

$\therefore$ pH = 7

Student X : $[H^+]$ or $[OH^-] = \frac{(0.2\times20)+(0.4\times20)-(0.1\times20)}{20+20+20+40}$

$\therefore$ $[H^+] = \frac{10}{100}$M

$\therefore$ pH = 1

Student Y : $[H^+]$ or $[OH^-] = \frac{(0.5\times14)+(0.2\times35)-(0.5\times8)}{14+35+8+43}$

$\therefore$ $[H^+] = \frac{10}{100}$M = 0.1

$\therefore$ pH = 1

Student Z : $[H^+]$ or $[OH^-] = \frac{(0.5\times20)+(0.4\times15)-(0.2\times30)}{20+15+30+35}$

$\therefore$ $[H^+] = \frac{10}{100}$M = 0.1 M

$\therefore$ pH = 1

20. Electron gain enthalpy is the energy change that occurs when an isolated gaseous atom gains an electron to form a negatively charged ion. It reflects the tendency of an atom to attract and bind an extra electron. Following experimental results arc provided in the table as shown below:

Process	Electron gain enthalpy ($kJmol^{-1}$)	Process	Electron gain enthalpy ($kJmol^{-1}$)
$O_{(g)} + e^- \rightarrow O^-_{(g)}$	-140	$S_{(g)} + e^- \rightarrow S^-_{(g)}$	-200
$O^-_{(g)} + e^- \rightarrow O^{2-}_{(g)}$	+780	$S^-_{(g)} + e^- \rightarrow S^{2-}_{(g)}$	+640

Choose the correct statement(s):

(A) First electron gain process is more favourable for sulphur than that of oxygen because electron-electron repulsion is more in the valence shell of oxygen atom than that of sulphur atom.

(B) Second electron gain process is more favourable for oxygen than that of sulphur because the size of O^{2-} anion is less than the size of S^{2-}.

(C) $O_{(g)} + 2e^- \longrightarrow O^{2-}_{(g)}$ process is more favourable than $S_{(g)} + 2e^- \longrightarrow S^{2-}_{(g)}$ because the first ionisation energy of neon is higher than the first ionisation energy of argon.

(D) In general, the second electron gain enthalpy is positive because a negatively charged anion resists accepting an additional electron more than a neutral atom does.

Ans. (A, D)

Sol. Release of energy is favourable for a reaction to occurs.

21. A salt mixture is made with 0.03 mole each of iron (II) oxalate, iron (III) oxalate, iron (II) sulphite, iron (III) sulphite, iron (II) nitrate and iron (III) nitrate in acidic medium. The mixture is oxidised separately with potassium permanganate and potassium diehromatc.

Find the CORRECT statements) relating to above oxidation processes.

(A) 33.53 g potassium dichromate is required for the oxidation process and two ions present in the salt mixture will not participate in the oxidation process.

(B) 27.95 g potassium dichromate is required for the oxidation process and three ions present in the salt mixture will participate in the oxidation process.

(C) 15.01 g potassium permanganate is required for the oxidation process and three ions present in the salt mixture will participate in the oxidation process.

(D) 18.02 g potassium permanganate is required for the oxidation process and two ions present in the salt mixture will not participate in the oxidation process.

Ans. (B, D)

Sol. $FeC_2O_4 \longrightarrow Fe^{3+} + CO_2$

$n = 3$

$Fe_2(C_2O_4)_3 \longrightarrow Fe^{3+} + CO_2$

$n = 6$

$FeSO_3 \longrightarrow Fe^{3+} + SO_4^{2-}$

$n = 3$

$Fe_2(SO_3)_3 \longrightarrow Fe^{3+} + SO_4^{2-}$

$n = 6$

$Fe(NO_3)_2 \longrightarrow Fe^{3+} + NO_3^-$

$n = 1$

$Fe(NO_3)_3 \longrightarrow$ No reaction

$K_2Cr_2O_7 \longrightarrow Cr^{3+}$

n = 6

$KMnO_4 \longrightarrow Mn^{2+}$

n = 5

Equivalent of all ions of salt = $(0.03 \times 3) + (0.03 \times 6) + (0.03 \times 3) + (0.03 \times 6) + (0.03 \times 1)$

$= 0.57$

$\therefore$ Mole of $K_2Cr_2O_7$ used = $\frac{0.57}{6} = 0.095$

$\therefore$ Wt. of $K_2Cr_2O_7$ used = $0.095 \times 294 = 27.93$ gm

$\therefore$ Moles of $KMnO_4$ used = $\frac{0.57}{5} = 0.114$

$\therefore$ Wt. of $KMnO_4$ used = $0.114 \times 158 = 18.02$ gm

22. Four geometrical areas in the shapes of a parallelogram, a rhombus, a right-angled isosceles triangle and an equilateral triangle are taken. A point mass is placed on the corners of every area. A unit mass must be placed at a point within the boundary of each of these areas so that it remains in equilibrium (masses interact gravitationally). Ignore the gravitational attraction due to the earth while working out this problem. Identify the correct options.

(A) If the unit mass is kept at the centroid of each geometrical areas, it will be in equilibrium.

(B) In the parallelogram and in the rhombus, the unit mass will be in equilibrium if placed at the intersection of their diagonals, and for equilibrium in the equilateral triangle, it should be kept on its in-centre.

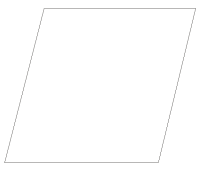
(C) In the rhombus and the parallelogram, the unit mass must be kept on the respective centroids, and in the equilateral triangle, the unit mass must be kept on its circumcentre.

(D) There is no point inside the right-angled isosceles triangle for keeping this unit mass so that it remains in equilibrium.

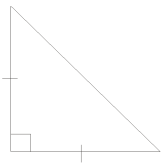
Ans. **(B, C)**

Sol. The gravitational field at the centroids of all the figures will be zero, except the isosceles right angle triangle.

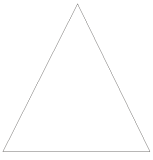
Parallelogram



Rhombus

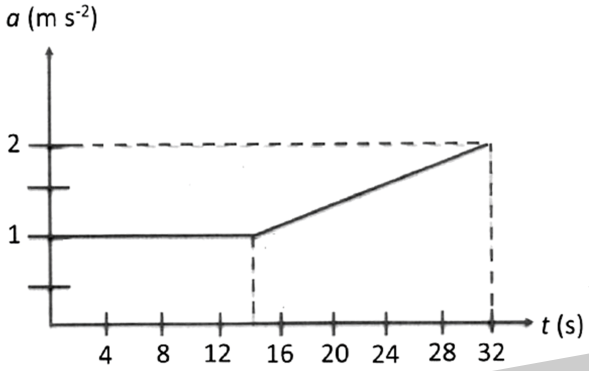


Isosceles right triangle



Equlateral triangle

23. In the graph shown below, the time-acceleration of a particle that has started from an initial velocity of 2 m s^{-1} is given. Identify the correct options given below.



(A) The panicle will have a displacement ot 657 m at the end ot 32 s.

(B) The displacement of the particle cannot be determined using average acceleration as the panicle is not moving with constant acceleration.

(C) The panicle will have a velocity 41 m s 1 after 32 s.

(D) The velocity of the particle at the end of 32 s is 43 m s^{-1}.

Ans. (B, D)

Sol. $v_0 = 2$ m/sec

$\Rightarrow$ Area under the entire graph

$$= 1 \times 14 + \frac{1}{2} \times 1 \times 18 + 1 \times 18 = 41 = v_{32} - v_0$$

$\Rightarrow v_{32} = 43$ m/s

$\Rightarrow$ Displacement from 0 – 14 sec

$$\Rightarrow S_1 = 2 \times 14 + \frac{1}{2} \times 1 \times 14^2 = 126 \text{ m}$$

$\Rightarrow$ Acceleration function from t = 14 – 32.

$$\Rightarrow a = \frac{1}{18}(t - 14) + 1$$

$$\Rightarrow a = \frac{t}{18} + \frac{2}{9} \text{ [Time starts from t = 14]}$$

$\Rightarrow$ After integration

$$\Rightarrow v = \frac{t^2}{36} + \frac{2}{9}t + c$$

$$\Rightarrow \text{at } t = 14,\ v = 16 \text{ m/s},\ c = \frac{67}{9}$$

$$\Rightarrow v = \frac{t^2}{36} + \frac{2}{9}t + \frac{67}{9} \quad ...(1)$$

Integrating the velocity expression.

$$\Rightarrow \int_0^{s_2}\left[\frac{t^3}{108}+\frac{t^2}{9}+\frac{67}{9}t\right]_{14}^{32}$$

$\Rightarrow$ $s_2 = 504$ m

$\Rightarrow$ Total displacement = $s_1 + s_2 = 126 + 504 = 630$ m

24. On an experiment board, four resistors, each of R = 100 Ω rated 25 W, are connected to a DC supply of the emf equal to E volt and negligible internal resistance, as shown in the figure. A fuse wire of negligible resistance has been connected for protection.

Choose the correct optipn/options.

(A) The maximum possible value of £ that can be applied, keeping the components safe is E = 125 V.

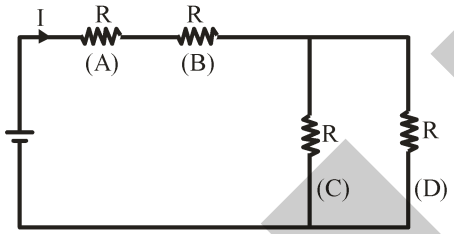
(B) The maximum power that can be dissipated in all four resistors together is 62.5 W.

(C) The current rating of fuse in the circuit is /max = 1 A.

(D) The total power dissipated in all four resistors together when E = 50 V is 10 W.

Ans. (A, B, D)

Sol.



Maximum current will flow through resistor A and B.

So $P_A = 25$

$$I^2R = \left[\frac{E}{\left(\frac{5R}{2}\right)}\right]^2 \times R = 25$$

E = 125 volt.

Total power of circuit = $\frac{E^2}{\frac{5R}{2}} = 62.5$ W

it E = 50 v then $\rightarrow$ total power = $\frac{E^2}{\frac{5R}{2}} = 10$ W

SECTION-III

25. A decade ago, Suresh, Neema, Gulzar, David, Moriri, Indali, and Udit were schoolmates in Delhi, united by their shared passion for long-distance running. Moriri and Neema hailed from the highlands of Himachal Pradesh and Arunachal Pradesh, respectively, while Indali and Gulzar were bom in coastal towns - Mangaluru and Vapi. Surcsh, Udit, and David, on the other hand, were bom in the plains: Chhattisgarh, Patna, and Anantapur respectively.

Fast forward to present; now they are spread across various parts of India, each following different career paths. Suresh now supervises operations at the Mumbai dockyard, while Moriri works with an IT company in Bengaluru. Indali, a sports representative for Indian Railways, continues her passion for athletics. Gulzar leads a team in a textile industry, and Udit is part of an R&D team in a biotcch company in Hyderabad. David serves in the Indian Army and is currently stationed in Ladakh. while Neema is a science communicator based in Bhopal.

David planned for a hiking and trekking adventure to reunite with friends at Leh region after 6 months during summer. All happily agreed upon and Neema reached a month ahead via a roadway and took this opportunity to trek throughout Leh-Ladakh, interacting and spreading awareness among locals on issues of global warming. (8 marks)

1. During summer, rest others reached Leh by flight. Within a weeks' time some of them started to haye mild to severe headache. Write the name(s) of people who had headache. **(2 Marks)**

2. Write the names of people who are relatively having higher RBCs in blood circulation than their other friends **(2 Marks)**

3. If all of them start to trek within a weeks' time who all will be having higher lactic acid production in muscles? **(2 Marks)**

4. Doctor said that these headaches are outcome of _____. To have relief from this headache, locals suggested them to take ginger tea. which causes _____. Fill in the blanks with appropriate reasons. **(2 Marks)**

Sol. **Hint : Native places**

Moriri, Neema = H.P, A.P.
Indali, Gulzar = Costal towns
Suresh, Udit, David – Plains.

Work places :

Suresh = Mumbai
Mariri = Bengaluru
Indali = Sports representitive
Gulzar, Udit = Hyderabad
David = Ladhkh
Neema = Bhopal

1. Headache = Reported in Suresh, Gulzar, Udit.
2. High R.B.Cs = Moriri, Neema, David, Indali
3. Gulzar
4. Doctor said that these headache are outcome of altitude sickness. To have relief from this headache, locals suggested them to take ginger tea. which increase level of testosterone which causes increase the R.B.Cs. production

26. There are four sub-questions. There is no negative marking for any sub-questions. **(22 marks)**

1. Transpiration takes place predominantly by stomata. They are specialised apparatus in epidermis with a pore space surrounded by two specialised cells called guard cells. These cells function as multisensory hydraulic valve. Environmental factors like light, temperature, relative humidity, and intracellular CO_2 are sensed by these guard cells and response leads to opening and closing of stomata. Mark the statement below as true and false with respect to stomatal movements. **(6 marks)**

(A) Blue light activates proton pumps in guard cells. It will lead to ion uptake and water influx, which opens the stomata.

(B) During stress conditions levels of abscisic acid increases. It will trigger opening of stomata to allow CO2 intake to increase photosynthetic activity.

(C) To open stomata, guard cells reduce turgidity and shrink itself reversibly.

(D) Proton pump present in membrane of guard cells facilitates potassium ion uptake.

(E) Potassium ions movement play key role and are actively transported into the guard cell during the opening and out during closing of stomata.

(F) Low CO_2 concentration in leaves triggers stomatal opening meeting the requirement for photosynthesis

2. Graft is a widely used horticultural technique for plant propagation and development of new plant varieties using two segments (Scion and Stock) of plants from two different resources. **(2 marks)**

Arrange the statements given below that describes the sequential steps leading to a successful graft formation.

(A) Formation of plasmodesmata at the cell-cell interface between scion and stock.

(B) Regeneration of the cortex and epidermis at the interface between scion and stocks

(C) Coordinated development of vascular tissues at the interface between scion and stocks.

(D) The cells at interface of scion and stock region go through a rapid phase of division forming callus.

3. Medicinal plants generally exhibit their therapeutic properties through the secondary metabolites. These are compounds that are produced by plants for various functions like defenee against pathogens and pests, attractants for pollinators, signaling etc.

Below is a list of some examples that are attributed to the functional effects of seconday metabolites of various plants and their parts.

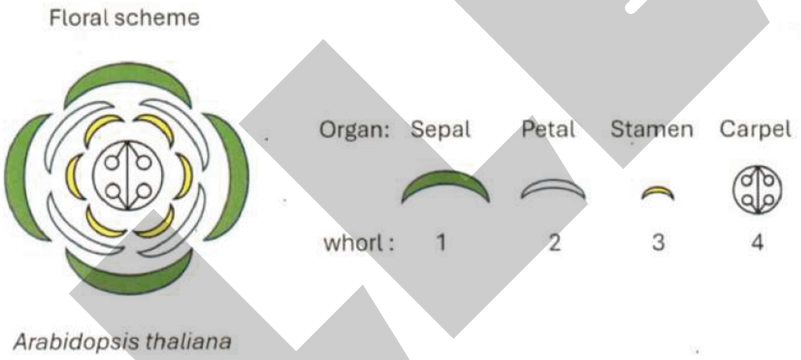
- Phenolics in the seeds consumed have antioxidant properties.

- European travelers to Asia and Africa were advised to take bark infusion water as a precaution against malaria.
- Phytotoxic secretions from their roots are known to inhibit growth of other plant around it.
- Oil extract of leaves is used traditionally for wound healing.
- The astringent effect of the seeds relieves tooth ache.

Match the secondary metabolites of column A with the used (in Column B) and the plant source (in Column C). **(5 marks)**

Column A	Column B	Column C
Secondary	Example	Plant source
1. Resveratrol	A. Anti-malarial	P. Areca catechu
2. Quinine	B. Analgesic	Q. Mangifera indica
3. Azadirachtin	C. Anti-inflammatory	R. Arachis hypogaea
4. Mangiferin	D. Insecticidal	S. Azadirachta indica
5. Catechins	E. Antioxidants	T. Cinchona pubsecens

4. In Arabidopsis thaliana, a small weed, is an excellent model to study role of genes instructing the flowering pattern and location of organs. Flowers of A. thaliana are made up of four organs: sepal, petal, stamen and carpel which are arranged m circular layers called whorls.



Scientist discovered that the development of the e l1ower parts is controlled by three groups of special genes called the A, B, and gen which act alone or in combination to decide which part or the l1ower grow in each whorl. When they are mutated individually, in all the instances the number of whorls remained the same (4), but the floral parts within the whorls replaced with flora parts from another whorl. It's never a mixture of random organs is any whorl.

- Mutation of gene A (Mutant A) resulted in flowers having only stamens and carpels.
- Mutation of gene B (Mutant B) resulted in flowers having only sepals and carpels.
- Mutation of gene C (Mutant C) resulted in flowers having only sepal and petals.

From the above information, deduce the outcome of the mutation(s) with respect to expression of organs in each whorl and fill in the complete word in each block. **(9 marks)**

Mutant type	Whorl Number			
	1	2	3	4
Wild Type	Sepals	Petals	Stamen	Carpels
Mutant A				
Mutant B				

Mutant C				
Mutant AB				
Mutant AC				
Mutant BC				
Mutant ABC	Leaf-like	Leaf-like	Leaf-like	Leaf-like

Sol. **1.** A-T, B-F, C-F, D-T, E-F, F-T

2. D → A → C → B

3.

Column – A	Column – B	Column – C
1	B	R
2	A	T
3	D	S
4	E	Q
5	C	P

4.

Mutant type	**Whorl Number**			
	1	**2**	**3**	**4**
Wild Type	**Sepals**	**Petals**	**Stamen**	**Carpels**
Mutant A	Carpel	Stamen	Stamen	Carpel
Mutant B	Sepal	Sepal	Carpel	Carpel
Mutant C	Sepal	Petal	Petal	Sepal
Mutant AB	Carpel	Carpel	Carpel	Carpel
Mutant AC	Leaf like	Leaf like	Leaf like	Leaf like
Mutant BC	Sepal	Sepal	Sepal	Sepal
Mutant ABC	Leaf-like	Leaf-like	Leaf-like	Leaf-like

27. **pH TITRATION:** Amit and Sumit have taken a joint school project under the guidance of their subject teacher to perform a weak acid (P) -strong base (Q) titration by using a pH meter and its theoretical verification at room temperature (30°C). Amit prepared an aqueous solution of Qin a 500 mL volumetric flask. 33 ml of this solution was utilised for complete neutralisation of 25 mL of a standard 1% (w/v) dibasic acid (molecular weight= 126). In another 500 mL volumetric flask, Sumit made an aqueous solution of 3.06% (w/v) monobasic acid P ($K_a = 4 \times 10^6$). 25 ml of 0.12 M standard KOH solution was required to neutralise 20 ml of solution of P.

Now, Sumit started titration of acid (P) with base (Q) with a pH meter. l 0 mL of aqueous acidic (P) solution is taken in a beaker and gradually small quantities of aqueous solution of base (Q) is added from the burette up to 19 ml mark. He noted the readings obtained • from pH meter.

Their guide-teacher discussed the following concepts and equations for calculations of pH theoretically.

$HA(aq) + H_2O(\ell) \rightleftharpoons H_3O^+(aq) + A^-(aq)$ (1)

$$K = \frac{[H_3O^+][A^-]}{[HA][H_2O]} \quad (2)$$

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} \quad (3)$$

$$pH = pK_a + \log \frac{[C_{salt}]}{[C_{acid}]} \quad (4)$$

$$pH = 7 + \frac{1}{2} pK_a + \frac{1}{2} \log C_{salt} \quad (5)$$

The ionization equilibrium of the acid P is represented by equation (1). Applying the laws of chemical equilibrium we get equation (2). Since, concentration of water is very large and remains almost constant in solution; it can be combined with K to give another constant K_a, the ionization constant of acid, represented by equation (3). Before neutralisation or equivalence point and after addition of Q, equation (4) is valid. And at equivalence point equation (5) should be used. Using these equations, Amit calculated pH of the solution mixture theoretically after each addition of Q in small quantities with above concepts and equations. Now, their guide-teacher compared both reading obtained from pH meter and theoretical calculation. **(21 marks)**

1. What is the molecular weight of monobasic acid P? **(2 marks)**

2. **(6 marks)**

2.1 What is the molarity of solution Q?

2.2 Complete the following table in the answer sheet:

Neutralisation of 10 mL of aqueous solution of P with aqueous solution of Q

SI. No.	Volume of aqueous solution of Q added (mL)	Calculated pH
1	0.0	
2	1.0	
3	12.0	
4	12.5	
5	13.0	
6	19.0	

3. **(5 marks)**

3.1 Draw a graph of volume (mL) of Q added vs calculated pH, by connecting the points as a smooth curve. (use the supplied graph sheet)

3.2 Label the discrete parts that you notice on the curve. (on the graph)

3.3 Show the pH at equivalence on the graph.

4. Provide a one/two sentence explanation for the transition that is noticed in terms of the concentration of the species in the solution, provide chemical equations to support your arguments. **(4 marks)**

4.1 below pH ≈ 4.5 and

4.2 between pH ≈ 6.5 and pH ≈ 11.5.

5. Considering the nature of the acid and base being used with respect to their ionisation power, justify the pH obtained at the equivalence point. Provide necessary chemical equations to support. **(2 marks)**

6. Which indicator(s) would be preferred for the above titration? Choose from the given table and provide a I-sentence explanation for your choice. **(2 marks)**

Indicator	Colours		Colour transition pH range
	Acid	Base	
Methyl violet	Yellow	Blue	0.0-1.6
Methyl orange	Red	Yellow	3.2-4.4
Methyl red	Yellow	Red	4.8-6.0
Bromothymol blue	Yellow	Blue	6.0-7.6
Phenol red	Yellow	Pink	6.8-8.0
Thymol blue	Yellow	Blue	8.0-9.6
Phenolphthalein	Colourless	Pink	8.2-10.0
Thymolphthalein	Colourless	Blue	9.3-10.5
Alizarin yellow	Yellow	Orange/Red	10.2-12.0

Sol. Experiment by Amit:

Molarity of disbasic acid solution = $\frac{1}{126} \times 10$

During neutralisation of solution Q by solution of dibasic acid

$$M \times 33 = \frac{10}{126} \times 2 \times 25$$

$\therefore$ Molarity of basic solution (Q) = 0.12 M

Experiment by Sumit:

During neutralisation of solution P by solution of KOH

$$\frac{3.06}{M} \times 10 \times 20 = 0.12 \times 25$$

(1) $\therefore$ Molar mass of acid P = 204 gm/mole

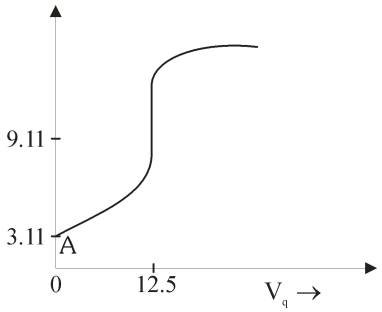
(2) (2.1) : Molarity of solution Q = 0.12 M

Molarity of acid solution P = $\frac{3.06}{204} \times 10 = 0.15$ M

(2.2) :

S.No.	Volume of aqueous solution of Q added	pH
1	0.0	3.11
2	1.0	4.34
3.	12.0	6.78
4	12.5	9.11
5.	13	11.42
6.	19	12.43

(3) (3.1)



(3.2) When neutralisation is about to complete, variation of pH of resulting solution is sharp.

(4) (4.1) : $HA_{(aq)} + BOH_{(aq)} \longrightarrow AB_{(aq)} + H_2O_{(\ell)}$

Till 1 ml of base added, since solution is still acidic so variation of pH is slow and below 7.

(4.2) : When 12 ml base solution is added.

$$\underset{(1.5-0.12)}{HA} + BOH \longrightarrow \underset{(0.12\ mm)}{BA} + H_2O$$

$$pH = pKa + \log\left(\frac{0.12}{1.5-0.12}\right) = 6.78$$

When 12.5 ml base solution is added

$$\underset{(1.5-1.5)\ mm}{HA} + BOH \longrightarrow \underset{(1.5\ mm)}{BA} + H_2O$$

$$\therefore\ pH = 7 + \frac{1}{2}\ pKa + \frac{1}{2}\ \log C$$

$$\therefore\ pH = 9.11$$

When 13 ml base solution is added

$$HA + \underset{(0.12\times 13-1.5)}{BOH} \longrightarrow \underset{(1.5)}{BA} + H_2O$$

$$\therefore\ [OH^-] = \frac{0.06}{23} = 0.0026\ M$$

$$\therefore\ pH = 11.42$$

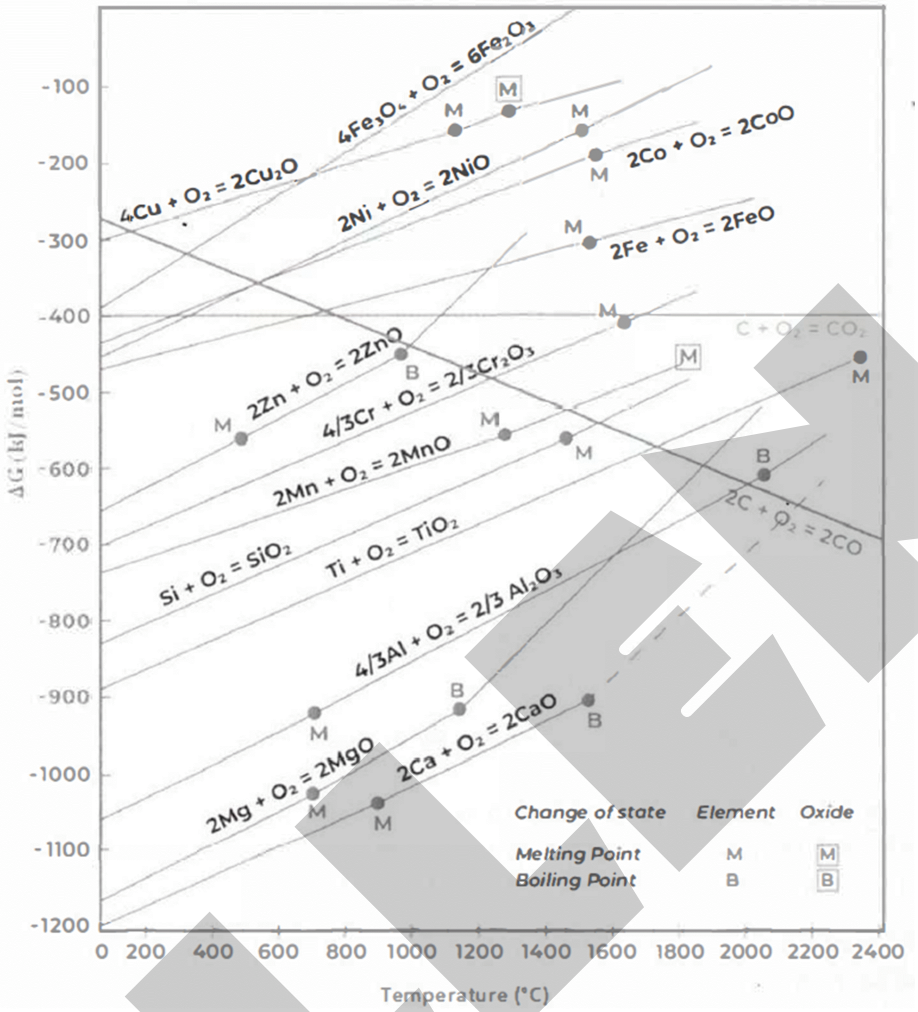
(5) Since salt formed at equivalence point is formed by weak acid and strong base. So, nature of resulting solution is basic.

(6) Thymol blue, phenolphthalein, Thymolphthalein can be used for above titration because their colour transition pH range lie in between sharp variation of pH when neutralisation is about to complete.

28. METALLURGY: Metals are extracted from their ores by various processes. Commonly, the metals in different chemical forms are converted into their oxides, and then these metal "Oxides are usually reduced to pure metal by using Carbon as reducing agent. Carbon combines with the oxygen of the metal oxide.

$$M_xO_y + yC \rightarrow xM + yCO$$

Ellingham diagrams are graphical representations used in metallurgy to predict the temperature dependence of the stability of oxides and other compounds. These diagrams plot the Gibbs free energy change (ΔG) of a reaction, typically involving the formation of a metal oxide after consuming one mole of oxygen, against temperature. Named after Harold Ellingham, who introduced them in 1944, they are valuable tools for understanding the thermodynamics of metal extraction and oxidation processes.



In an Ellingham diagram. the y-axis represents the Gibbs free energy change (ΔG) for the formation of oxide shown as metal + oxygen gas, and the x-axis represents temperature, usually in degrees Celsius or Kelvin. The change in Gibbs enery, ΔG for any process at. any specified temperature, is described by the equation:

$$\Delta G = \Delta H - T\Delta S$$

where, ΔH is the enthalpy change and ΔS is the entropy change for the process. Enthalpy is a measure of the total heat content of a system. Entropy is a measure of the randomness or disorder of a system. Entropy increases from solids to liquids· to gases due to the increasing freedom of movement and arrangement of molecules.

Ellingham diagrams are useful in determining which reducing agent can be employed to reduce a particular metal oxide. A reducing agent wil1 only reduce an oxide if it results in a decrease in Gibbs free energy. On the diagram, this can be visualized by comparing the position of the lines for the metal oxide and the reducing agent's reaction with oxygen.

The reducing agent must have a more negative ΔG at the specific temperature for the reduction to be spontaneous. Therefore. the magnitude of the difference in the ΔG values for the formation of oxides helps in selecting the reducing agent. The greater the difference, the more spontaneous the reaction.

For example, carbon is often used as a reducing agent in metallurgy. The line for the formation of carbon monoxide from carbon intersects with the lines for various metal oxides at different temperatures. This intersection point indicates the temperature above which carbon can reduce that specific metal oxide to its metal form. The Ellingham diagram thus helps metallurgists determine the optimal temperature and conditions for smelting and refining metals, such as in the extraction of iron in a blast furnace. **(9 marks)**

Based on the information given about Ellingham diagrams, answer the following:

1. Explain based on the nature of the reactants and products, why most of the lines in the Ellingham diagram slope upwards. **(2 marks)**

2. Carbon forms two different oxides. Lines corresponding to formation of one of the oxides slopes downward with increase in temperature and the other one stays horizontal. Give a one/two sentence explanation for this phenomenon. **(2 marks)**

3. Bharathi was working in the research section of a metallurgical industry. She received an oxide ore each of Zinc and chromium to determine the conditions to extract the metals from their respective ores. She referred the Ellingham diagram. What would be the minimum temperature at which carbon can be used as a reducing agent to reduce the zinc ore? Write a balanced chemical equation denoting the process. **(2 marks)**

4. What are the possible metals that can reduce CriO3? Arrange them in the increasing order of their reducing ability at 800 °C. Which among them would be the most economical reducing agent that can be used. **(3 marks)**

Sol. 1. Since, during oxidation ΔS of reaction is negative So, the slope of lines is positive.

2. During formation of CO_2 gas from carbon, mole of gases remain constant so ΔS is zero.

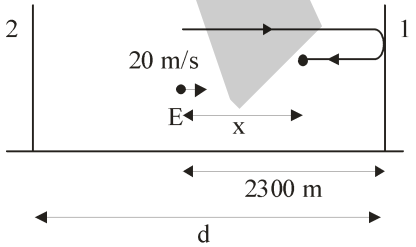
3. $ZnO_{(s)} + C_{(s)} \longrightarrow Zn_{(s)} + CO_{(g)}$

The above reaction occur spontaneously above 1000°C.

4. Cr_2O_3 can be reduced by Mn, Si, Ti, Al, Mg, Ca. in the increasing order of their reducing ability. Most economical reducing among them is Ca.

29. An engine is running at a speed of 72 km hr- 1 towards cliff 1 along a railway track laid between two cliffs 1 & 2 separated by d km. The engine makes a· whistle when it is 2.3 km away from the cliff in front of it. The driver of the engine receives the first echo from cliff l after 13s. If the engine d1iver receives the 2nd echo 4 s after the first one. fnd the distance between the two cliffs and the velocity of sound. When will the engine driver receive a tbjrd echo, and what direction would that come from? **(5 marks)**

Sol.

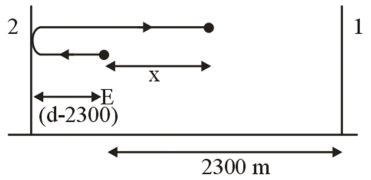


$x = 20 \times 13 = 260$ m

Distance travelled by sound during

13 sec $\rightarrow$ 2300 + (2300 – 260) = 4340 m

Speed of sound = $\frac{4340}{13}$ = 333.84 m/s



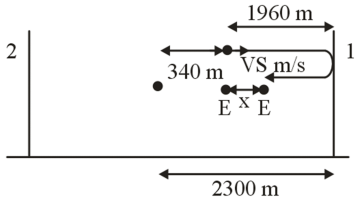
For 2nd echo → x = v_E × 17 = 20 × 17 = 340 m

Distance travelled by sound = 333.84 × 17 = 2 (d – 2300) + 340

⇒ d = 4967.64 m

d ≈ 4.9 km

For third echo →



x = v_E × t = 20 t

x + v_st = 2 × 1960

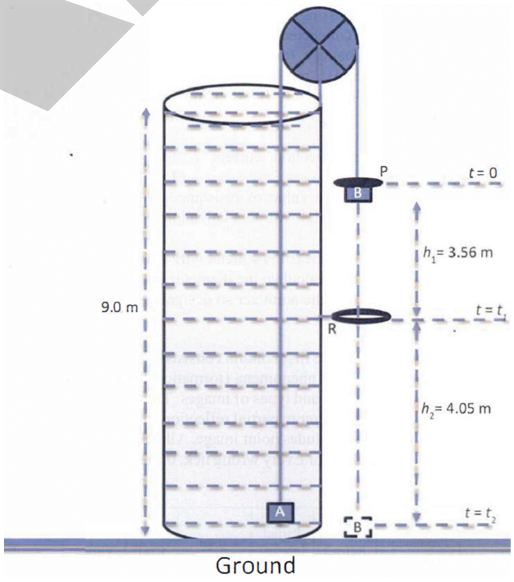
20 t + (333.84) t = 3920

(353.84) t = 3920

t = $\frac{3920}{353.84}$ = 11 sec.

So time of third echo from staring = 17 + 11 = 28 sec.

30. A tank placed on the ground is filled with water up to a height of 9.00 m. A frictionless light pulley is attached to the top edge of the tank. Two identical cubical blocks. A and B. each of edge length IO cm and relative density 5, are connected to two free ends of a light inextensible string which passes over this pulley. Thus. block A is hanging inside the water and block B is in the air. A rider P of mass I kg is placed on block B. The arrangement is shown in the figure given below.



When the system is released at instant t = 0, block B with the rider descends through a distance h1 for time t_1 At this instant, the rider is caught by the· ring R that allows block B only to pass through it. At latter instant of time t_2, the block strikes the ground. Assume that air resistance is negligible, and the effects of surface tension and viscosity of water are ignored. Density of water is 1000 kg m-3. Answer the following questions. **(12 marks)**

1. What are the masses of block A and B? **(1 marks)**

2. Calculate the acceleration of block A and tension in the string during time t = 0 to t_1. **(3 marks)**

3. Calculate the acceleration of block Band tension in the string during time t = t_1 to t_2. **(3 marks)**

4. Find out the values of t_1 and t_2. **(3 marks)**

5. Calculate the velocities of block A at instants of time t_1 and t_2. **(2 marks)**

Sol. 30. density of blocks = 5000 kg/m^3

(a) masses of blocks A and B

$= 5000 \times (10 \times 10^{-2})^3$

$= 5$ kg

(b) Acceleration of system :

$\Rightarrow a = \dfrac{6g - 5g + (P_L Vg)}{11}$

$\Rightarrow P_L = 1000, V = (10 \times 10^{-2})^3, g = 10$

$\Rightarrow a_1 = \dfrac{20}{11}$ m/s^2 = 1.81 m/s^2

Tension in strong = T

$\Rightarrow 6g - T = 6 \times \dfrac{20}{11}, T = \dfrac{540}{11}$ N

$\Rightarrow T = 49.09$ N

(c) $\Rightarrow a_2 = \dfrac{P_L Vg}{10} = 1$ m/s^2

$\Rightarrow 5g - T = 5 \times 1$

$\Rightarrow T = 45$ N

(d) $3.56 = \dfrac{1}{2} \times \left(\dfrac{20}{11}\right) \times t_1^2$

$\Rightarrow t_1 = 1.978$ sec

$\Rightarrow v_1 = 0 + \left(\dfrac{20}{11}\right) \times t_1 = 3.597$

$\Rightarrow v_2^2 - v_1^2 = 2 \times 1 \times 4.05$

$\Rightarrow v_2 = 4.586$

$\Rightarrow v_2 - v_1 = a_2 \times (t_2 - t_1)$

$\Rightarrow t_2 - t_1 = 0.989$

$\Rightarrow t_2 = 2.967$ sec.

(e) $v_1 = 3.597$ m/s

$v_2 = 4.586$ m/s

31. A moving coil galvanometer of the resistance R_G = 50 Ω has equally spaced 30 division graduated on its circular scale. The deflection produced in galvanometer is directly proportional to the current passing through it.

When this galvanometer is connected in a single loop closed circuit in series with a battery (emf = 2.0 V, internal resistance is negligible) and a resistor R = 9950 Ω, the current produces a deflection of 10 division on the galvanometer scale. **(5 marks)**

1. Calculate fult scale deflection current I_g for this galvanometer.

2. Calculate the approximate value of resistance S that must be connected in parallel with this galvanometer to measure the current range 0 to 3A. **(1.5 marks)**

3. If the shunt resistance in part (b) is a metal wire of length 15.7 cm of material having resistivity 2×10^{-8} Ωm, calculate the diameter of the wire. **(1.5 marks)**

4. Estimate the resistance of the ammeter so designed. **(1 mark)**

Sol. Current in galvanometer for reading of 10 divisions →

$$I = \frac{2}{9950+50} = \frac{2}{10,000} A$$

1. $I_g = 3I = \frac{6}{10,000} A$

2. $I_g R_g = (3 - i_g)s$

$$\frac{6}{10,000} \times 50 = \left(3 - \frac{6}{10,000}\right) s$$

$s = 0.01\ \Omega$

3. $R = \rho \ell / A$

$$0.01 = 2 \times 10^{-8} \times \frac{15.7 \times 10^{-2}}{\pi r^2}$$

$r^2 = 9.99 \times 10^{-8}$

$r = 3.16 \times 10^{-4}$ m

$D = 2r = 6.32 \times 10^{-4}$ m

4. $R_A = \frac{R_g \times s}{R_g + s} = \frac{50 \times 0.01}{50 + 0.01}$

$= 0.0099\ \Omega$

32. Consider the following grid. The first column consists of optical objects (like mirrors, glass lenses, etc., kept in the air) and phenomena (formation of rainbow, mirage). The first row gives a list of optical prope1ties and types of images. The incident beam is monochromatic. It can be diverging or parallel. Ignore partial reflection from the first surface of glass lenses or prism. Diminished image includes point image. All you should do is to place tick mark/s (✓) in all the applicable cell/s. For Every wrong tick, 0.25 mark will be deducted. **(8 marks)**

	Prominent reflection	Refraction	Partial Internal reflection	Total internal reflection	Real image	Virtual image	Image size greater than object size	Image size smaller that object size	Image of same size
Metallic Convex mirror	✓					✓		✓	
Metallic Concave mirror	✓				✓	✓	✓	✓	✓
Equilateral glass prism		✓	✓	✓		✓			✓
Thin biconvex lens		✓	✓	✓	✓	✓	✓	✓	✓
Prism binoculars		✓	✓	✓		✓	✓		
Metallic plane mirror	✓					✓			✓
Rainbow		✓	✓	✓	✓		✓	✓	✓

ALLEN'S RESULT IN IJSO

(LAST 10 YEARS)

SESSION	INDIA's MEDAL			ALLEN's MEDAL		
	GOLD	SILVER	BRONZE	GOLD	SILVER	BRONZE
2014-15	5	1	–	2	1	–
2015-16	5	1	–	4	–	–
2016-17	4	2	–	4	2	–
2017-18	5	1	–	4	1	–
2018-19	6	–	–	6	–	–
2020-21	6	–	–	4	–	–
2021-22	6	–	–	5	–	–
2022-23	5	1	–	5	1	–
2023-24	6	–	–	4	–	–
TOTAL	**48**	**6**	**–**	**38**	**5**	**–**

Total Medals of INDIA 54

Total Medals of ALLEN 43

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ONCE AGAIN **ALLEN** STUDENTS BRING

GLORY TO THE NATION

5 GOLD & 1 SILVER MEDAL AT IJSO 2023

IJSO Thailand 2023

20th International Junior Science Olympiad

Left to Right

Kanishk Jain (Workshop) - Gold Medal | **Mahroof Khan** (Classroom) - Gold Medal | **Archit Bhalodiya** (Classroom) - Silver Medal
Soham Pednekar (Classroom) - Gold Medal | **Rudra Pethani** (Classroom) - Gold Medal | **Divya Agrawal** (Classroom) - Gold Medal

This is for the 3rd time,
Team ALLEN is Team INDIA at IJSO World Stage

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ALLEN

Gold Medals Unlocked!

4 ALLENites strike gold at IJSO 2024

IJSO Romania 2024

21st International
Junior Science Olympiad

Left to Right

HARSHIT SINGLA
(Classroom)

JINANSH J. SHAH
(Classroom)

PRANIT MATHUR
(Classroom)

MANAS GOEL
(Classroom)

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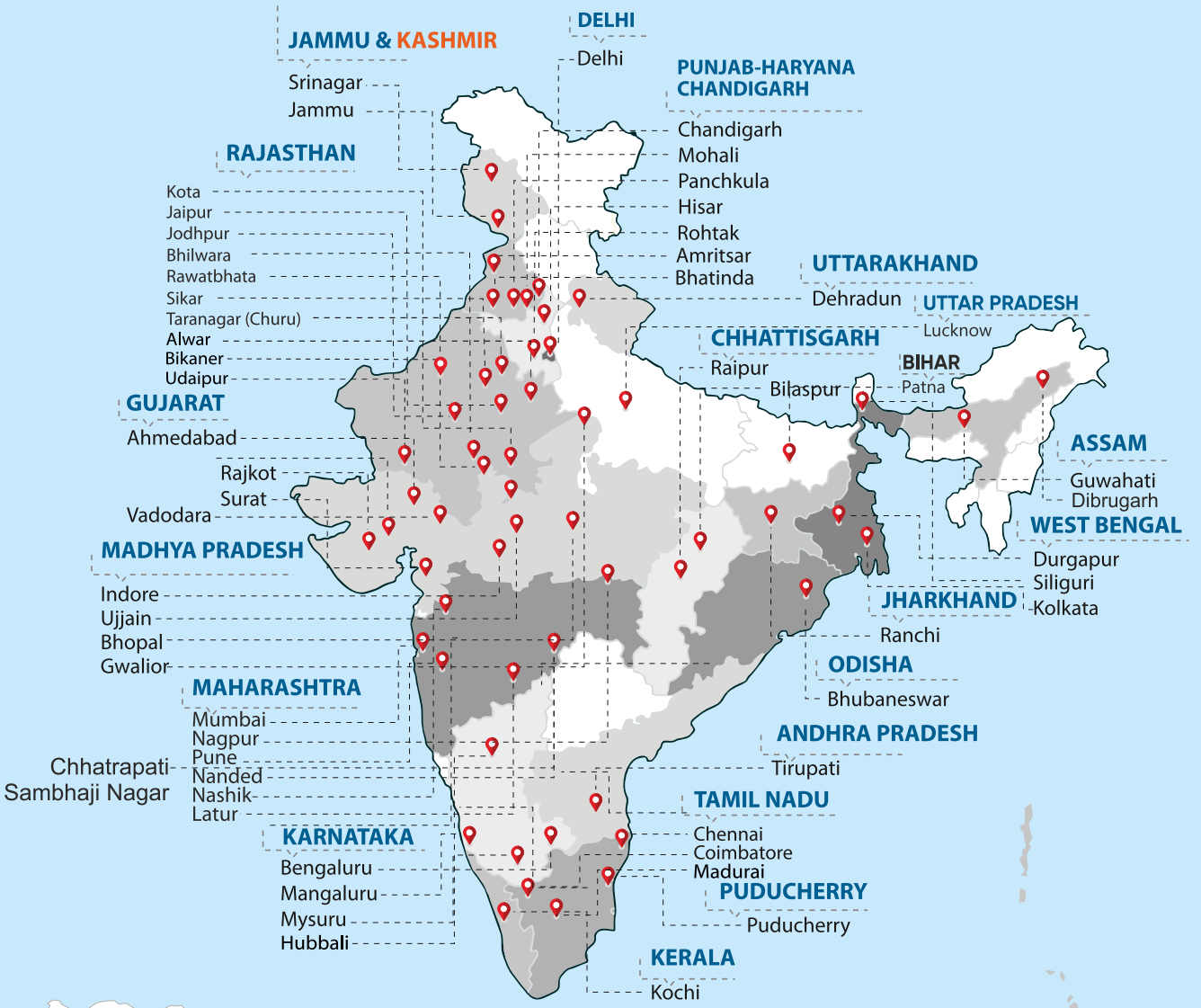

MAP NOT TO SCALE

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- QATAR
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